**Volumetric Redox A.O.L.**

1. Convert the following to PPM

a) 0.75g/l = \_\_\_\_\_\_\_\_\_\_\_\_\_\_

b) 0.0003g/100cm3= \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2. Calculate in grams the following:

a) 1M NaOH = \_\_\_\_\_\_\_\_\_\_\_\_

b) 0.125 M H2SO4= \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

3. What is the concentration in moles per Litre of the following solutions

a) 1 g of H2SO4 in 1 litre = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

b) 24g NaOH in 500cm3 = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

c) 4 g HCL in 750cm3= \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

d) 5%w/v of NaOH=\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

e) 0.2 %w/v H2SO4 =\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

4. Calculate the number of moles of sulfuric acid (H2SO4) in:

a) 50 cm3 of 0.1 M H2SO4

b) 100cm3 of 0.05 M H2SO4

5.

